Practice

Strings of p orbitals
- move π bonds towards π bonds
- move π bonds towards empty p orbitals.
- move lp e⁻'s towards π bonds
- move lp e⁻'s towards empty p orbitals

doesn't work
- can't put a π bond here because of the 2H atoms
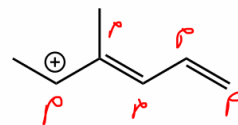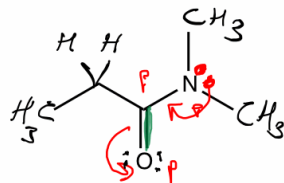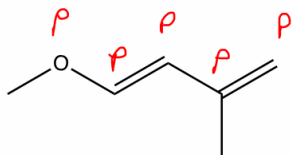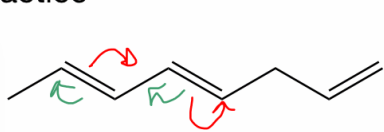- can't change the charge of the molecule

Practice
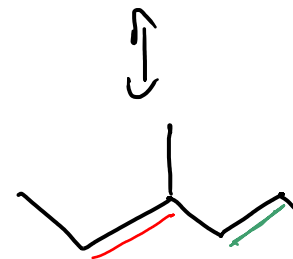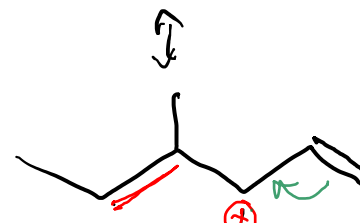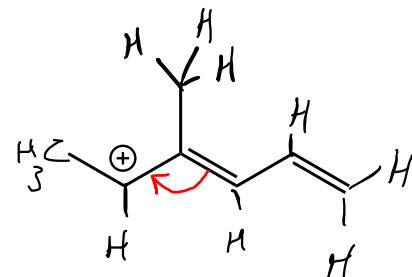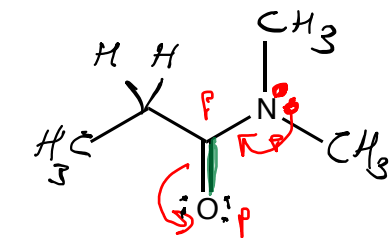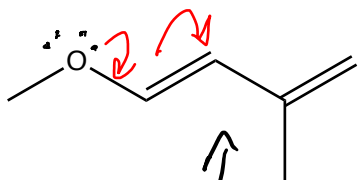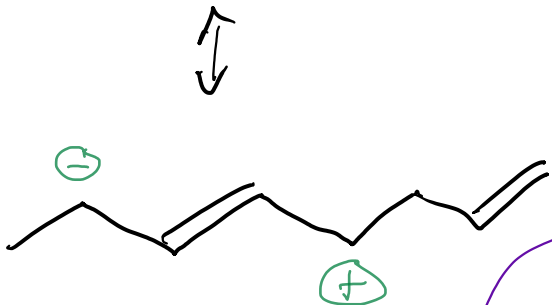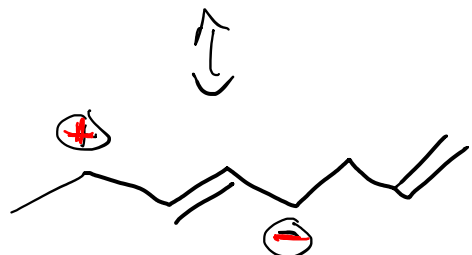
most important
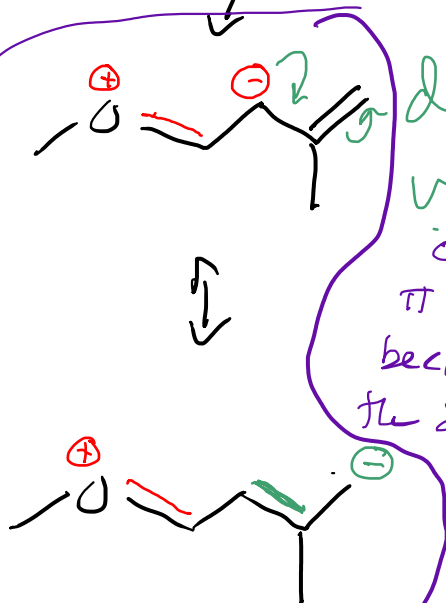
g doesn't work

'can't put a
π bond here
because of
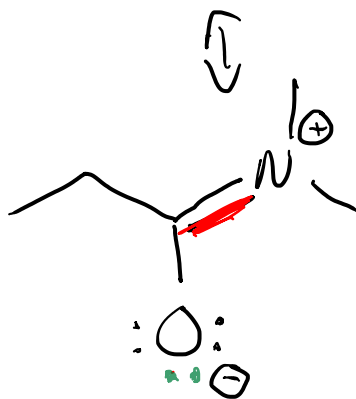the 2H atoms

cant
change the charge
of the molecule

too
many
bonds

2+ charges

charge separation takes energy to do that
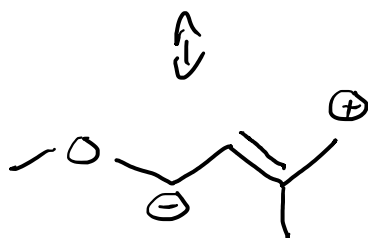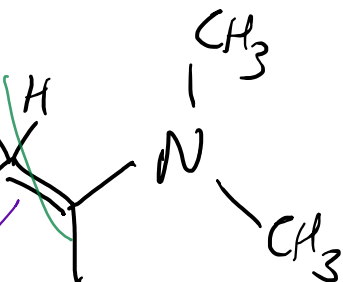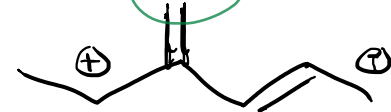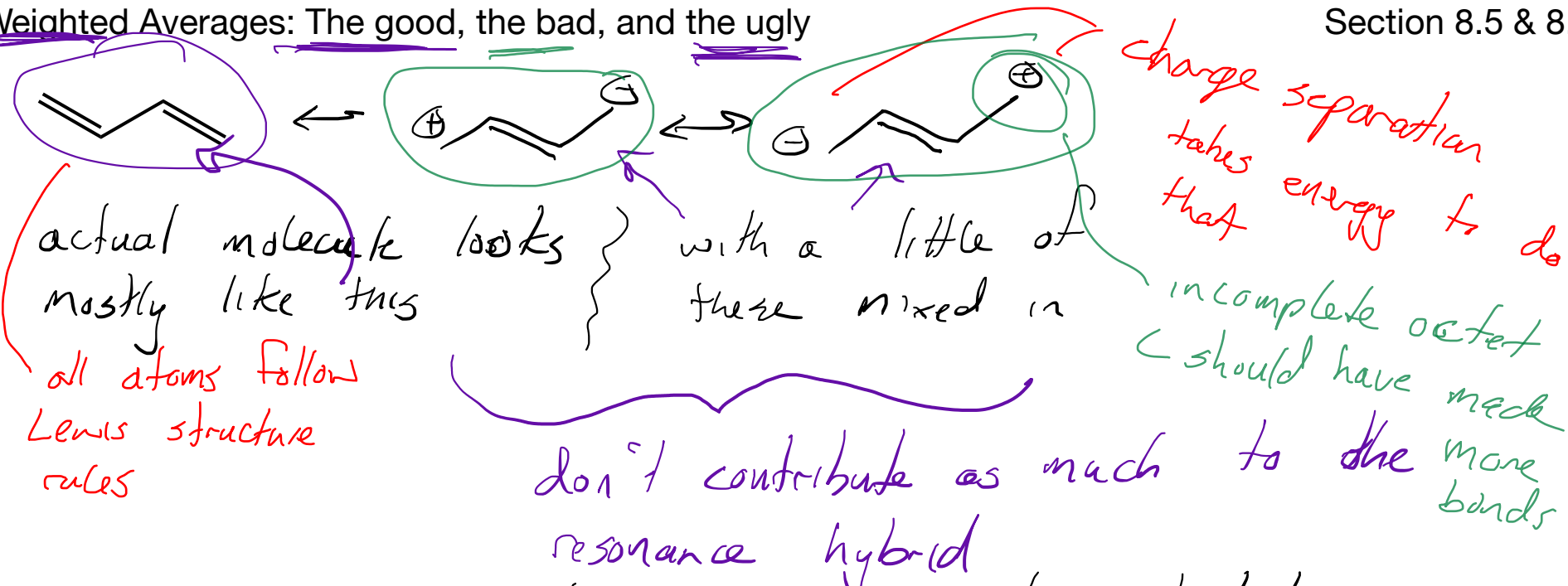
incomplete octet ← should have made more bonds

actual molecule looks } with a little of these mixed in

all atoms follow Lewis structure rules

don't contribute as much to the resonance hybrid

The "better" the contributor the more it contributes to the resonance hybrid.

Creating charge separation

Creating atoms with incomplete octets }

Putting the "wrong" charge on an atom

least eneg atom with ⊖

most eneg atom with ⊕

all three problems would create a structure that would be an insignificant contributer to the resonance hybrid don't draw it